

- ①  $2\text{Al}(s) + 3\text{S}(s) \rightarrow \text{Al}_2\text{S}_3(s)$  synthesis
- ②  $\text{H}_2\text{O}(l) + \text{N}_2\text{O}_5(g) \rightarrow 2\text{HNO}_3(aq)$  synthesis
- ③  $4\text{NO}_2(g) + \text{O}_2(g) \rightarrow 2\text{N}_2\text{O}_5(g)$  synthesis + combustion
- ④  $2\text{C}_2\text{H}_6(g) + 7\text{O}_2(g) \rightarrow 4\text{CO}_2(g) + 6\text{H}_2\text{O}(g)$  combustion
- ⑤  $2\text{Al}_2\text{O}_3(s) \rightarrow 4\text{Al}(s) + 3\text{O}_2(g)$
- ⑥  $\text{Ni}(\text{OH})_2(s) \rightarrow \text{NiO}(s) + \text{H}_2\text{O}(l)$
- ⑦  $2\text{NaHCO}_3(s) \rightarrow \text{Na}_2\text{CO}_3(s) + \text{CO}_2(g) + \text{H}_2\text{O}(l)$
- ⑧  $2\text{K}(s) + 2\text{HCl}(aq) \rightarrow 2\text{K}(s) + 2\text{KCl}(aq)$
- ⑨ NR
- ⑩ NR
- ⑪  $\text{LiI}(aq) + \text{AgNO}_3(aq) \rightarrow \text{AgI}(s) + \text{LiNO}_3(aq)$
- ⑫  $\text{BaCl}_2(aq) + \text{K}_2\text{CO}_3(aq) \rightarrow \text{BaCO}_3(s) + 2\text{KCl}(aq)$
- ⑬  $\text{Na}_2\text{C}_2\text{O}_4(aq) + \text{Pb}(\text{NO}_3)_2(aq) \rightarrow \text{PbC}_2\text{O}_4(s) + 2\text{NaNO}_3(aq)$